

Field Testing for Ozone



Grades: 6-12

Subject: Earth Science, Chemistry

Time: 2 Class Periods (1 Week Apart)

Student Objectives

- Distinguish between tropospheric and stratospheric ozone.
- Define smog.
- Discuss air quality and related health and welfare issues involving tropospheric and stratospheric ozone.
- Perform an experiment that demonstrates that ozone can have a detrimental effect on certain materials.

Materials

- Potassium iodide
- Filter paper (can use coffee filters)
- Corn starch
- Glass stirring rod
- Small paint brush
- 250 mL beaker
- Heat source (preferably a hot plate)
- 9" microwaveable plate or paper plate
- Zipper lock plastic bags (~ 4)
- Distilled water
- Heat safe glass plate
- Full-splash safety goggles
- Aprons
- Scissors
- Schoenbein Color Scale (provided)
- Relative Humidity Schoenbein Number Chart (provided)
- Bulb psychrometer (optional)

Student Activity

This test is based on the oxidation capability of ozone. Ozone in the air will oxidize the potassium iodide on the Schoenbein paper to produce iodine. The iodine reacts with starch and produces a purple color. The exact shade of purple correlates to the amount of ozone present in the air. The two reactions involved are:



I_2 + starch \rightarrow starch turns a shade of purple

Background Information

The issue of ozone in the earth's atmosphere can be confusing. On one hand, we know that high above the earth's surface in the stratosphere is a layer of ozone that surrounds the planet and helps block out some of the sun's harmful radiation. We hear reports of "holes" developing in this stratospheric ozone shield and of the harm that the increased ultraviolet radiation can cause on earth. On the other hand, we know that higher than normal concentrations of ozone in the air we breathe in the troposphere can be harmful to people, animals, plants, and various materials. The ozone gas in the stratosphere and troposphere is the same, the chemical O_3 . In the upper atmosphere (stratosphere) it greatly benefits all life. Near the earth's surface (troposphere), it can cause problems.

Background Info (Cont.)

The Stratospheric Ozone Layer

High in the stratosphere a layer of ozone gas forms an important and effective protective barrier against the harmful ultraviolet radiation from the sun. There has been increasing international concern that chemical pollutants are destroying this ozone layer. The main culprits seem to be a class of chemical compounds called chlorofluorocarbons, or CFCs. Since the early 1970's, researchers across the globe have been working to understand the ozone layer, the causes of its depletion, and the effects of ozone layer depletion on humans and the environment. Visit EPA's website to learn more:

<https://www.epa.gov/ozone-layer-protection>.

Increased ultraviolet radiation at the earth's surface can lead to a greater incidence of:

- skin cancer,
- eye problems,
- immune deficiencies in humans,
- decreased crop yields,
- reduced populations of microscopic sea plants, and
- animals that are vital to the food chain.

Ozone Pollution in the Troposphere

High concentrations of ozone in the ambient air that we breathe in the troposphere can present many problems. Because ozone molecules are highly reactive, they have an effect on practically every material they contact, whether it be lung tissue, crops or other vegetation, rubber, plastic, paints, etc.

What we often refer to as photochemical "smog" is mostly ground-level ozone. The recipe for the formation of ozone in the ambient air includes volatile organic compounds (VOCs), nitrogen oxides, and sunlight. Because sunlight is a key factor, ozone pollution is generally worse during the day and in the summertime. Vehicle exhaust provide most of the VOCs and nitrogen oxides that help form ozone, so times of increased vehicle use (such as morning and afternoon rush hours) also increase the possibility of ozone problems.

Ozone can cause eye, nose, and throat irritation, and can damage the lungs. Visit EPA's website to learn more about about ozone pollution in the troposphere: <https://www.epa.gov/ozone-pollution>.

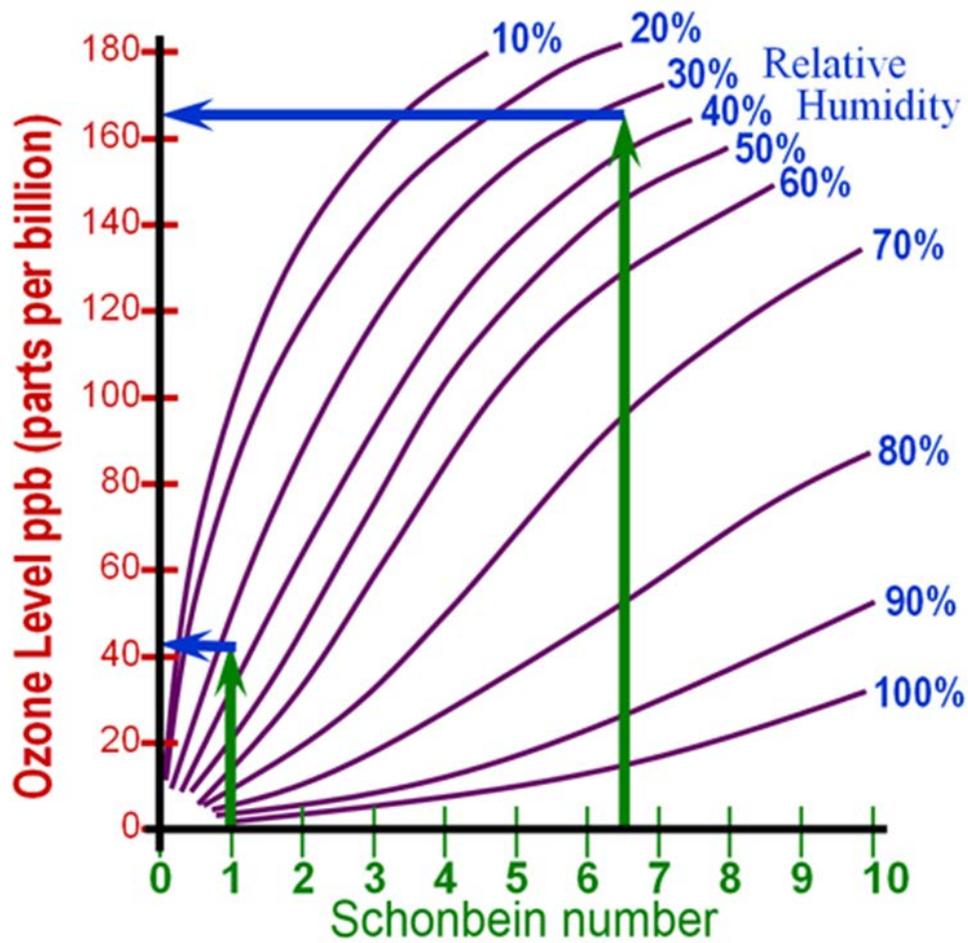
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Student Directions

1. Place 100 mL of water in a 250 mL beaker on your heat source, then add 5 grams of corn starch.
2. Heat and stir mixture until it gels. The mixture is gelled when it thickens and becomes somewhat translucent.
3. Remove the beaker from the heat and add 1 gram of potassium iodide and stir well. Allow the solution to cool.
4. Lay a piece of filter paper (you can use coffee filters) on a glass plate and carefully brush the paste onto the filter paper. Turn the filter paper over and do the same on the other side. Apply the paste as uniformly as possible. The paper can be exposed for immediate testing at this point.
5. Allow the paper to dry. To save time, place the paper on a microwave-safe plate and microwave for one minute.
6. Cut the paper into small strips. To store the strips, place them in a zipper lock plastic bag or glass jar out of direct sunlight.
7. Dip a strip of test paper in distilled water and hang it at a data collection site out of direct sunlight. Make sure the strip can hang freely.
8. Expose the paper for approximately eight hours. Seal it in an airtight container if the results will not be recorded immediately.
9. To observe and record test results, dip the paper in distilled water. Observe the color and determine the Schoenbein Number using the Schoenbein color scale.
10. Determine the relative humidity of the data collection site by using a bulb psychrometer or local weather data.
11. Round off the relative humidity reading to the nearest 10 percent. (Higher relative humidity makes the paper more sensitive to ozone, and a higher Schoenbein Number is observed. To correct for this, the relative humidity must be determined and figured into the calculation of the ozone concentration.)
12. Refer to the Relative Humidity Schoenbein Number Chart. Along the bottom of the chart, find the point that corresponds to the Schoenbein number that you recorded. From that point, draw a line upward until it intersects with the curve that corresponds to your relative humidity reading. To find the ozone concentration in parts per billion, draw a perpendicular line from the Schoenbein number/relative humidity point of intersection to the left side of the chart (see example below).
NOTE: The color of the paper may not be uniform. Determine the Schoenbein Number by the color in the area with the most noticeable change.

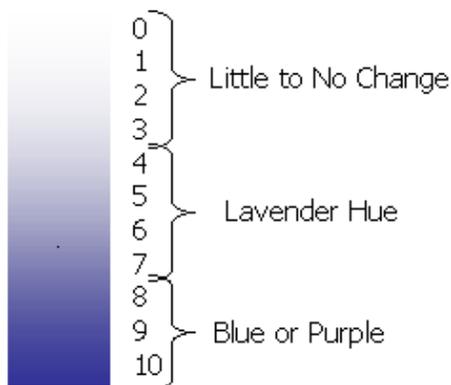
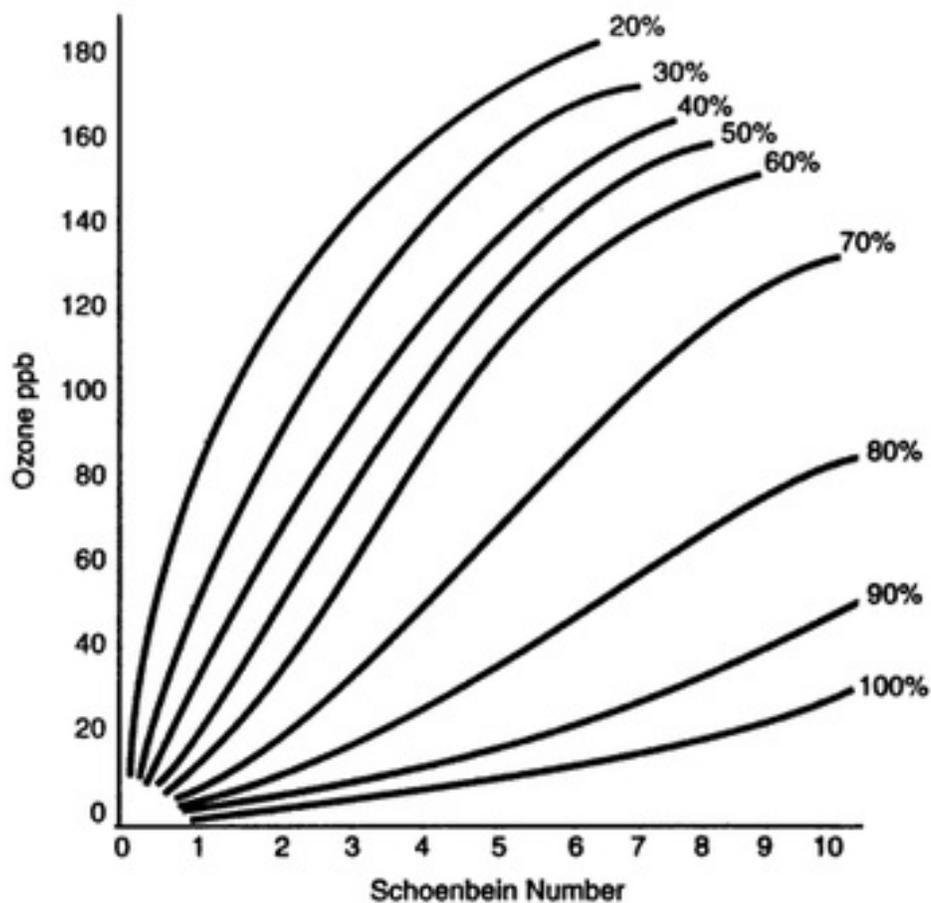
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Relative Humidity Schoenbein Number Chart



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Observations & Questions

1. What change in the test paper, if any, did you observe?
2. Compare your test paper to those of other students. Do all the test papers appear the same? (Individual test papers will vary depending on the amount of oxidants at that site. Be aware that false positive results can occur from nitrous oxides in heavy traffic areas.)
3. Was the relative humidity for your test day high or low? (Individual results will vary depending on the specific relative humidity of the site.)
4. Why do you think the test papers did not all appear the same?
5. Would the ozone parts per billion (ppb) be the same for a Schoenbein Number of 4 at a relative humidity of 30 percent and 70 percent? (Hint: Refer to the Relative Humidity Schoenbein Number Chart.)
6. Based on the data you collected, do you think this method is a good way to measure tropospheric ozone? Why or why not?
7. Compare data with those from a local monitoring station. Also, if possible get information about the wind direction during your study and determine how it affected your measurements.

What Can Be Done?

Both ozone problems, stratospheric depletion and tropospheric build-up, are created in large part by air pollution. The only practical approach to stopping the destruction of the ozone layer and to minimizing ozone pollution in our ambient air is reducing the human-generated pollutants that contribute to these problems. Finding and using alternatives to CFCs is an essential part of the solution. As individuals we can immediately repair any leaks in refrigerators, have our car air conditioners checked periodically, use alternatives to home air conditioning, use alternatives to foam insulation and containers, purchase halon-free fire extinguishers, and support laws requiring CFC recycling.

Decreasing our use of vehicles burning fossil fuels and assuring our vehicle emission control systems are functioning properly is also critical to solving the problem of tropospheric ozone. We can use public transportation for long trips, walk or use bicycles for short trips, carpool to work and other activities, and combine several errands into one outing. Some areas have "ozone action" days, which encourage citizens and industries to follow procedures to reduce their impact on the formation of harmful ozone. On these days, citizens are encouraged to postpone mowing their lawns and refilling their automobile's gas tanks until the evening hours, avoid using lighter fluid for charcoal, and carpool or use public transportation.

Add link where they can find data.